

Lecture notes

Lukas Rahmn

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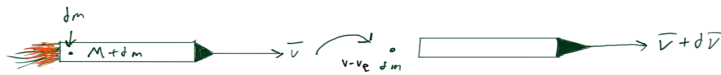
Lite repetition
raket ekv
värmelära
Mekanik problem

Problemlösnings metoder

1. Newtons 2:a lag, a
2. Mek energi, v
3. \vec{P} bevaras

repetition

Raketekvationen



$$P_i = (M + dm)V$$

$$P_f = M(v + dv) + dm(v - v_e)$$

$$P_i = P_f, \text{ Inga externa krafter}$$

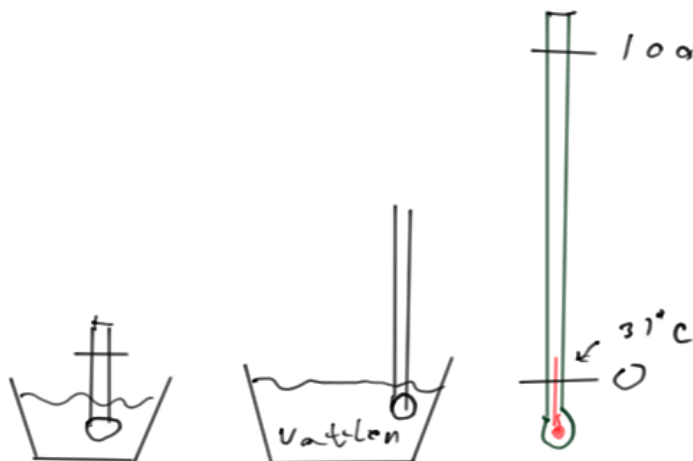
$$(M + dm)v = M(v + dv) + dm(v - v_e)$$

$$\rightarrow Mdv = dm v_e, dm = -dM \rightarrow Mdv = -dM v_e$$

$$\rightarrow \int_{v_i}^{v_f} dv = -v_e \int_{M_i}^{M_f} \frac{dM}{M} \rightarrow v_f - v_i = v_e \ln \left(\frac{M_i}{M_f} \right)$$

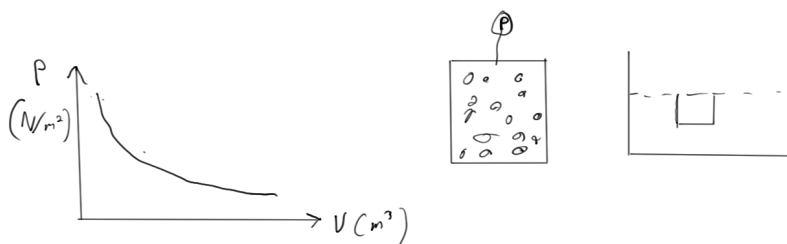
Värmelära

Temperatur



Dålig för höjd över havet, och utvidgningen av vätskan påverkar

Gasttermometer



$$PV = \text{konst}T$$

Vid 1 mol gas: konst = 8,31

$$PV = nRT, R = 8.31 \text{ J/K}, n = \text{Antal mol}$$

$$\text{Avrogadostal } N_A = 6,023 \cdot 10^{23} \text{ st}$$

Enheter för tryck

$$1. 1 \text{ M/m}^2 = 1 \text{ Pa}$$

$$2. 1 \text{ bar} = 10^5 \text{ N/m}^2$$

$$3. 1 \text{ atm} = 1,013 \cdot 10^5 \text{ N/m}^2$$

1dim:

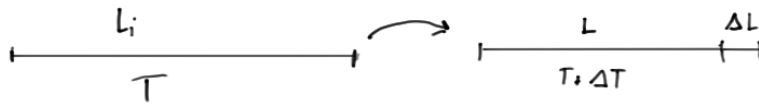
Millimeter kvicksilver:

$$P(Ah)\rho g/A$$

$$P = \rho gh$$

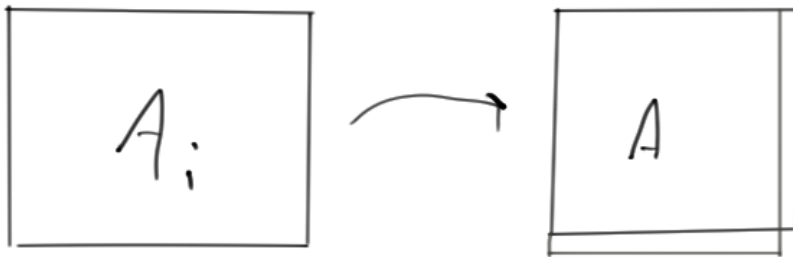
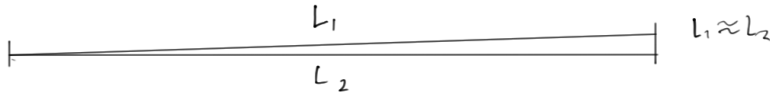
$$13,6 \cdot 10^3 * 0,700 * 9.81 = 1.013 \cdot 10^5$$

Termisk utvidning



$$\Delta L = L_i \alpha \Delta T$$

α : linjära längdutvecklingskoeff, typiskt värde: $\alpha = 10^{-6}$



2 dim:

$$A = A_i + \Delta A, \Delta A = A_i \beta \Delta T.$$

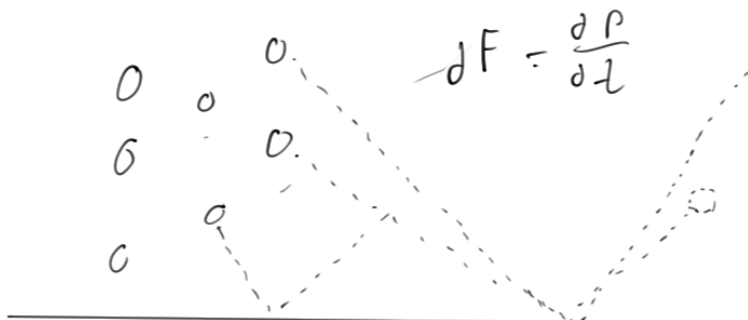
Tillbaka till ideala gaslagen

$$PV = nRT, \quad \begin{array}{|c|} \hline n \text{ mol} \\ \hline N \\ \hline \end{array}$$

$$PV = \frac{N}{N_A} RTP = \left(\frac{N}{V}\right) \left(\frac{R}{N_A}\right) T \rightarrow P = \left(\frac{N}{V}\right) k_B T$$

Boltzmanns konstant: $k_B = 1,38 \cdot 10^{-23} \text{ J/K}$

Vad är egentligen temperatur?



Medelenergin ges av

$$E_{\text{kinetisk energi}}^{\text{medel}} = \frac{3}{2}k_B T$$

$\frac{1}{2}k_B T$ i medelenergi per frihetsgrad

Värme – inre energi - Arbete

Inre energi E^{int}

$$\begin{aligned} E^{\text{int}} &= N \left(\frac{3}{2} k_B T \right) = \\ &= N \frac{3}{2} \frac{R}{N_A} T = \frac{N}{N_A} \left(\frac{3}{2} R \right) T \rightarrow E^{\text{int}} = \\ &= n \left(\frac{3}{2} R \right) T \end{aligned}$$

Enatomiga molekyler (He, Ne, Ar, ...) har 3 frihetsgrader ($E_{\text{medel}} = \frac{3}{2}k_B T$).
Tvåatomiga molekyler (H_2, O_2, N_2, \dots) har 5 frihetsgrader, $E_{\text{int}} = \frac{5}{2}k_B T \rightarrow E_{\text{int}} = n \left(\frac{5}{2} R \right) T$

$$dE^{\text{int}} = n \left(\frac{3}{2} R \right) dT$$